

Download Calculating Ph Of A Salt Solution

Calculating the pH of a Salt Solution. To calculate the pH of a salt solution one needs to know the concentration of the salt solution, whether the salt is an acidic, basic, or neutral salt, the equation for the interaction of the ion with the water, the equilibrium expression for this interaction and the K_a or K_b value. The pH of a solution of ammonium chloride can be found in a very similar way to the sodium fluoride solution in Sample Problem 21.7. However, since the ammonium chloride is acting as an acid, it is necessary to know the K_a of NH_4^+ , which is 5.6×10^{-10} . We will find the pH of a 2.00 M solution of NH_4Cl . In the case of salt solution $C_a = C_b$, but we have a good reason to treat these values separately. Once the general equation will be derived we can always simplify it using only one concentration, but general equation will describe not only pH of the salt solution, but also of any mixture of weak acid and weak base, regardless of their ratio. Ancient Rome Did NOT Build THIS Part 2 - World's LARGEST Stone Columns - Lost Technology - Baalbek - Duration: 9:51. Bright Insight 887,385 views